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Concept Of Chemistry
03 : MOLARITY and
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 371–372 1. Interpret
 the following balanced
 chemical equations in
 terms of particles,
 moles, and mass. Show
 that the law of
 conservation of mass is
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 Key CHAPTER 9
 REVIEW
 ...Stoichiometry
 Chapter 11 Study
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 Section 11.3, for
 example, you learned
 how to express the
 stoichiometry of the
 reaction for the
 ammonium dichromate
 volcano in terms of the
 atoms, ions, or
 molecules involved and
 the numbers of moles,
 grams, and formula
 units of each
 (recognizing, for
 instance, that 1 mol of
 ammonium dichromate
 produces 4 mol of
 water). Chapter 11.4:
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 STOICHIOMETRY. MOLE
 TO MOLE RATIO. When
 nitrogen and hydrogen
 gas are heated under
 the correct conditions,
 ammonia gas (NH_3) is
 formed. a. RXN: $1. \text{N}_2$
 $+ 3. \text{H}_2 \rightarrow 2. \text{NH}_3$. b.
 How many moles of
 nitrogen react with
 three moles of
 hydrogen? ___ 1 mol
 N_2 ___ 3 mol H_2 1 mol
 N_2 . 3 mol H_2 . c. How

many moles of nitrogen react with six moles of hydrogen?

 2 mol N₂ 6

mol H₂ 1 mol N₂

...CHAPTER 11:

STOICHIOMETRY Stoichi

ometry The study of

quantitative

relationships between

the amounts of

reactants used and

amounts of products

formed by a chemical

reaction is called

stoichiometry.

Stoichiometry is based

on the law of

conservation of mass.

Recall that the law

states that matter

is neither created nor

destroyed in a

chemical

reaction. CHAPTER 11

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metry is the tool for

answering these

questions.

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the amounts of

reactants used and

amounts of products

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reaction is called

stoichiometry.

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study of quantitative relationships between the amounts of reactants used and amounts of products formed by a chemical reaction is called stoichiometry. Stoichiometry is based on the law of conservation of mass. Recall that the law states that matter is neither created nor destroyed in a chemical reaction.

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amount of product that is expected to be produced in a chemical reaction (using stoichiometric calculations based upon the balanced chemical equation) and the amount of product actually produced when carried out in an actual experiment.

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Answer Key**

In Section 11.3, for example, you learned how to express the stoichiometry of the reaction for the ammonium dichromate volcano in terms of the atoms, ions, or molecules involved and the numbers of moles, grams, and formula units of each (recognizing, for instance, that 1 mol of ammonium dichromate produces 4 mol of

water).

*Chapter 11 Study
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Answers*

Stoichiometry is the tool for answering these questions. Stoichiometry The study of quantitative relationships between the amounts of reactants used and amounts of products formed by a chemical reaction is called stoichiometry.

Stoichiometry is based on the law of conservation of mass.

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of Chemistry
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[the amounts of.... In a](#)
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[ratio between the](#)
[numbers of moles.... A](#)
[reactant that remains](#)
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Problems pages
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gas are heated under
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formed. a. RXN: 1. N₂
+ 3. H₂ (2. NH₃. b.
How many moles of
nitrogen react with
three moles of
hydrogen? 1 mol
N₂ 3 mol H₂ 1 mol
N₂. 3 mol H₂. c. How
many moles of
nitrogen react with six
moles of hydrogen?
2 mol N₂ 6
mol H₂ 1 mol N₂ ...

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