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$V \times 25.0 \text{ mL OH}^- = 12.5 \text{ mmol HCl(aq)}$
 $\text{HCl(aq)} [\text{HCl(aq)}] 12.5 \text{ mmol} = n V$
 $0.10 \text{ mmol HCl(aq)} / \text{mL } V = 120 \text{ mL}$ 2. (a)
 Given: $V \text{ HCl(aq)} = 25.00 \text{ mL}$; $[\text{HCl(aq)}]$
 $= 0.350 \text{ mol/L}$; $V \text{ NaOH(aq)} = 5.0 \text{ mL}$;
 $[\text{NaOH(aq)}] = 0.500 \text{ mol/L}$; Required:
 $[\text{H}^+(\text{aq})]$, pH Analysis: Before titration,
 amount of H^+ is: $\text{H(aq)} \text{ H(aq)} \text{ H(aq)}$
 $0.350 \text{ mmol mL} + + + = x = n c V$ Section
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 Identify any errors in the structure by
 drawing them. Rename the structure
 correctly. 2-ethyl-2,4,4-trimethylpentane
 What Is Required? You are to identify the
 errors in the name by drawing the
 structure and renaming it correctly.

What Is Given? You are given an
 incorrectly named alkane. Unit 1 Organic
 Chemistry - Mr. Arthur's Science Page =
 0.12 mol/L Statement: The equilibrium
 concentration of dinitrogen tetroxide gas
 is 0.12 mol/L . 2. Given: initial quantity of
 $\text{NOCl(g)} = 3.00 \text{ mol}$; volume = 3.00 L ;
 $[\text{NO(g)}]$ initial = 0.00 mol/L ; $[\text{Cl}_2(\text{g})]$
 initial = 0.00 mol/L ; $[\text{NO(g)}]$ equilibrium
 $= 0.043 \text{ mol/L}$ Required: $[\text{NOCl(g)}]$
 equilibrium; $[\text{Cl}_2(\text{g})]$ equilibrium
 Solution: Step 1. Section 7.1: Equilibrium
 Systems Mini Investigation: The
 ... Solution: bonds broken $3 \text{ mol } 1 \text{ mol CH}_4$
 $\text{CCl}_3 \text{ mol HnD DD} -- \Delta = x = x + x = \sum$
 $413 \text{ kJ mol} \times +1 \text{ mol } 339 \text{ kJ mol} \times \Delta H$
 $= 1578 \text{ kJ}$ Statement: The energy
 required to separate 1 mol of
 chloromethane into free atoms is 1578
 kJ . 2. Solution: Step 1: Write the
 balanced chemical equation for the
 combustion of 1 mol of ethene gas. $\text{C}_2\text{H}_4(\text{g}) + 3 \text{ O}_2(\text{g})$... Section 5.3: Bond
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 Manual - Nelson Solution: $q = Ne$
 $= (1.2 \times 10^8)(1.602 \times 10^{-19} \text{ C})$
 $q = 1.92 \times 10^{11} \text{ C}$ (one extra digit carried)
 $F E = k q_1 q_2 / r^2 = 8.99 \times 10^9 \text{ N} \cdot \text{m}^2 / \text{C}^2$ #
 $\$ \% \& ' ((1.92 \times 10^{11} \text{ C}) (1.92 \times 10^{11} \text{ C})$
 $(1.10 \text{ m})^2 F E = 2.7 \times 10^{-12} \text{ N}$ Statement:
 The force of repulsion between the two
 plastic spheres is $2.7 \times 10^{-12} \text{ N}$. 2.
 Given: $m = 2.48 \times 10^{-15} \text{ kg}$; $\Delta d = 1.7$
 $\text{cm} = 0.017 \text{ m}$; $\Delta V_b = 260 \text{ V}$; $e = 1.602$
 $\times 10^{-19} \text{ C}$; $g = 9.8 \text{ m/s}^2$ Required: q ; N
 Solution: $\text{OH}^- \text{ OH}^- \text{ OH}^- 0.50 \text{ mmol mL} --$
 $-- = x = n c V \times 25.0 \text{ mL OH}^- = 12.5$
 $\text{mmol HCl(aq)} \text{ HCl(aq)} [\text{HCl(aq)}] 12.5$
 $\text{mmol} = n V 0.10 \text{ mmol HCl(aq)} / \text{mL } V$
 $= 120 \text{ mL}$ 2. (a) Given: $V \text{ HCl(aq)} = 25.00$
 mL ; $[\text{HCl(aq)}] = 0.350 \text{ mol/L}$; $V \text{ NaOH(aq)}$
 $= 5.0 \text{ mL}$; $[\text{NaOH(aq)}] = 0.500 \text{ mol/L}$;
 Required: $[\text{H}^+(\text{aq})]$, pH Analysis: Before

titration, amount of H^+ is: $H(aq) \rightarrow H^+(aq)$
 $H(aq) \quad 0.350 \text{ mmol} \quad mL \quad ++ \quad +=x = nc \quad V$

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current concept of organic chemistry is the study of carbon compounds. Each section of the chapter addresses a different group of organic compounds, related by structure. Throughout the chapter, tutorials provide a step-by-step procedure for naming and drawing structural diagrams. Each section also features important reactions of each compound ...

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Solution: bonds broken 3 mol 1 mol CH
 $CCl_3 \text{ mol} \quad HnD \quad DD \quad -- \quad \Delta = x = x + x = \Sigma$
 $413 \text{ kJ mol} \quad x \quad +1 \text{ mol} \quad 339 \text{ kJ mol} \quad x \quad \Delta H$
 $= 1578 \text{ kJ}$ Statement: The energy required to separate 1 mol of chloromethane into free atoms is 1578 kJ. 2. Solution: Step 1: Write the balanced chemical equation for the

combustion of 1 mol of ethene gas. $C_2H_4(g) + 3 O_2(g) \rightarrow 2 CO_2(g) + 2 H_2O(l)$...

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Solution: $q = Ne = (1.2 \times 10^8)(1.602 \times 10^{-19} \text{ C})$
 $q = 1.92 \times 10^{-11} \text{ C}$ (one extra digit carried)
 $F E = kq_1 q_2 / r^2 = 8.99 \times 10^9 \text{ N}^2 \text{ m}^2 \text{ C}^2 / \text{C}^2 \text{ m}^2$
 $F E = 2.7 \times 10^{-12} \text{ N}$
 Statement: The force of repulsion between the two plastic spheres is $2.7 \times 10^{-12} \text{ N}$.
 2. Given: $m = 2.48 \times 10^{-15} \text{ kg}$; $\Delta d = 1.7 \text{ cm} = 0.017 \text{ m}$; $\Delta V = 260 \text{ V}$; $e = 1.602 \times 10^{-19} \text{ C}$; $g = 9.8 \text{ m/s}^2$
 Required: q ; N

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True/False. Indicate whether the sentence or statement is true or false.
 1. The region in space where an electron is most likely to be found is called an energy level.
 2. Electron configurations are often condensed by writing them using the previous noble gas core as a starting point. ...

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$= 0.12 \text{ mol/L}$
 Statement: The equilibrium concentration of dinitrogen tetroxide gas is 0.12 mol/L .
 2. Given: initial quantity of $\text{NOCl(g)} = 3.00 \text{ mol}$; volume = 3.00 L ;
 $[\text{NO(g)}]_{\text{initial}} = 0.00 \text{ mol/L}$; $[\text{Cl}_2(\text{g})]_{\text{initial}} = 0.00 \text{ mol/L}$;
 $[\text{NO(g)}]_{\text{equilibrium}} = 0.043 \text{ mol/L}$
 Required: $[\text{NOCl(g)}]_{\text{equilibrium}}$; $[\text{Cl}_2(\text{g})]_{\text{equilibrium}}$
 Solution: Step 1.

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 8. Identify any errors in the structure by drawing them. Rename the structure correctly.
 2-ethyl-2,4,4-trimethylpentane
 What Is Required? You are to identify the errors in the name by drawing the structure and renaming it correctly.
 What Is Given? You are given an incorrectly named alkane.

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