
Chapter 9 Section 3

Stoichiometry

Answers

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Chapter 9 - Stoichiometry ~~Chapter 9 Section 3~~
Chapter 9 Section 3

9.1 Introduction to Stoichiometry ~~Chapter 3–~~
~~Chemical Reactions and Reaction Stoichiometry~~
~~Chapter 3– Stoichiometry and Calculations with~~
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130 - Session 17 - Chapter 9 Section 3 Chapter 3 -
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Equations: Part 4 of 5 ~~Chapter 3– Stoichiometry,~~
~~Formulas and Equations: Part 1 of 8 Step by Step~~
~~Stoichiometry Practice Problems | How to Pass~~
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~~Number Method~~ **Stoichiometry: What is**
Stoichiometry? *Stoichiometry (Mol Ratio)*

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Chapter 3 - Stoichiometry, Formulas and Equations: Part 4 of 8 *Stoichiometry Sample Problems Part 3* Chemistry (10-11) Chapter 9 sections 3 and 4.

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1- Identify the
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the reaction 3-
Set up the
problem
(convert to
moles if
necessary) 4-
Use the mole
ratio(s) to
calculate the
number of
moles of the
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<i>and</i>	<i>The Magic</i>	<i>Aqueous</i>
<i>Calculations</i>	<i>Number</i>	<i>Solution: Part</i>
<i>with Formulas</i>	<i>Method</i>	<i>1 of 8 Chapter</i>
<i>and</i>	Stoichiometry:	<i>4—Reactions</i>
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<i>and</i>	<i>Tutorial</i>	<i>Stoichiometry,</i>
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<i>Stoichiometry,</i>	<i>+ review</i>	<i>Chapter 3 -</i>
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<i>Equations:</i>	<i>explained †</i>	<i>and</i>
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<i>How to Pass</i>	<i>Reactant and</i>	<i>Gen Chem †</i>
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Reactions in Aqueous Solution: Part 5 of 8
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space provided. 1. 88% The actual yield of a reaction is 22 g and the theoretical yield is 25 g. Calculate the percentage yield. 2. 6.0 mol of N_2 are mixed with 12.0 mol of H_2 according to the following equation: $N_2(g) + 3H_2(g) \rightarrow 2NH_3(g)$

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as Fe₂(C₂O₄)₃ (s). Using a precisely measured tube called a burette, the student finds that 9.04 mL of the Na₂C₂O₄ solution was ...

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CHAPTER 9 REVIEW Stoichiometry SECTION 9-3 PROBLEMS

Write the answer on the line to the left. Show all your work in the space provided.

1. 88% If the actual yield of a reaction is 22 g and the

theoretical yield is 25 g, calculate the percent yield.

2. 6.0 mol of N₂ are mixed with

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FeCl₃. The student carefully adds 0.1074 M Na₂C₂O₄ until all the Fe³⁺

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Steps for Stoichiometry:

- 1- Identify the given and target compound
- 2- Balance the equation for the reaction
- 3- Set up the problem (convert to moles if necessary)
- 4- Use the mole ratio(s) to calculate the number of

moles of the desired compound 5-convert to grams of the desired compound if necessary

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